Homework
Class 9th A, B
Subject: Chemistry

Note: **Answer all the questions in your notebooks. Also learn these answers by heart.**

Recommended book: **Chemistry 9th Text book by Caravan Book House, Lahore.**

Assignment: 01	Chapter:5
	From 6th April to 19th April

1. What is diffusion (gases), explain with an example?
2. Define standard atmospheric pressure. What are its units? How it is related to Pascal?
3. Why are the densities of gases lower than that of liquids?
4. Gases are compressible. Why?
5. State Charle's law and write its mathematical derivation.
6. State Boyles law and write its mathematical derivation.
7. Define effusion and give an example.
8. What do you mean by evaporation how it is affected by surface area?
9. Define evaporation. How it is affected by intermolecular forces?
10. Differentiate between amorphous solids and crystalline solids.
11. Define the term allotropy with examples. Give two conditions of allotropy.
12. Convert the following units:
    (a) 850 mm Hg to atm (b) 205000 Pa to atm
    (c) 560 torr to cm Hg (d) 1.25 atm to Pa
    . Convert the following units:
    (a) 750 °C to K (b) 150 °C to K
    (c)100K to °C (d)172Kto°C.

Assignment:01	Chapter:6
	From 20th April to 26th April

1. Define solution and give an example.
2. Differentiate between solute and solvent.
3. What is an aqueous solution? Give example.
4. Define unsaturated and saturated solution.
5. What is a supersaturated solution?
6. Differentiate between dilute and concentrated solution.
7. Define concentration.
8. What is percentage - mass/mass (%m/m) ?
9. What is percentage - mass/volume (%m/v) ?
10. What is percentage - volume/mass (%v/m) ?

11. What is percentage - volume /volume (% v/v)?
12. Define molarity and write its formula.

Assignment:01	Chapter:6
	From 27th April to 2nd May

1. What is dilution of solution?
2. Describe preparation of one molar solution .
3. Define solubility.
4. What do you mean, like dissolves like? Explain with examples.
5. How does nature of attractive forces of solute-solute and solvent-solvent affect the solubility?
6. How you can explain the solute-solvent interaction to prepare a NaCl solution?
7. Justify with an example that solubility of a salt increases with the increase in temperature.

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