

Homework
Class 8th A, B
Subject: Chemistry

NOTE: Answer all questions in your notebook. Also learn these answers by heart.

CHAPTER:1

Fundamentals of Chemistry

- 1) Differentiate between physical and chemical properties with examples.
- 2) Differentiate between biochemistry and industrial chemistry.
- 3) Explain valency by giving suitable examples.
- 4) Write down valencies of following radicals/ elements:
Sodium, Potassium, Silver, Magnesium, Calcium, Barium, Zinc, Copper, Mercury, Iron, Aluminium, Chromium, Ammonium, Hydroxide, Bicarbonate, Hydrogen, Chlorine, Bromine, Iodine, Oxygen, Sulphur, Carbonate, Sulphate, Sulphite and Phosphate.
- 5) Write down formulas of following compounds.
Water, Common salt, Sand, Caustic soda, Washing soda, Quick lime, Lime stone, Sugar, Sulphuric acid and Ammonia
- 6) Define is relative atomic mass?
- 7) Write down various steps to write down chemical formulae of compounds.
- 8) Differentiate between empirical and molecular formula by giving examples.

- 9) What is relationship between empirical and molecular formula.
- 10) Differentiate between molecule and molecular ion.
- 11) Differentiate between ions and free radicals.
- 12) Define gram atom, gram molecule, gram formula and formula unit.

CHAPTER:2

Structure of atoms

- 1) How Electron was discovered?
- 2) How Proton was discovered?
- 3) How Neutron was discovered? Give its properties.
- 4) Explain salient features of Rutherford's atomic model.
- 5) State the postulates of Bohr' atomic theory.
- 6) What is plum pudding theory and who presented it?
- 7) Explain the uses of U-235 in power generation.
- 8) What is the number of protons and neutrons in an element having atomic mass 238 and atomic number 92?
- 9) Explain isotopes of carbon and chlorine.
- 10) What is meant by quantum and quanta. ?
- 11) A patient has goiter. How will it be detected?
- 12) For what purpose C-14 is used?
- 13) Describe how radiotherapy is used for treatment of cancer?

- 14) What is nature of charge on cathode rays?
- 15) How electrons are filled in various shells and subshells.

CHAPTER: 3

Periodic Table and periodicity of properties

- 1) Define Dobereiner's triads. Give its drawback.
- 2) Explain Newland's Octaves. Give its limitations.
- 3) What is Mendeleev's periodic law?
- 4) What are demerits of Mendeleev's periodic table?
- 5) State the Modern periodic law.
- 6) Write the salient features of long form periodic table.
- 7) What do you mean by groups and periods in the periodic table?
- 8) What do you mean by blocks in a periodic table and why elements were placed in blocks?
- 9) Discuss in detail the periods in periodic table.
- 10) What are transition elements?

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